Name: _____ Period: _____

Identifying Unknown Elements

Directions: Using a periodic table, identify the elements described in the statements below.

1. This element is in the same family as lead, and it has fewer protons than sodium. _____

2. This element has an atomic number that is one greater than platinum. _____

3. This element has the most protons of any element in Group 15. _____

4. This element has more than 50 but less than 75 protons, and it is in Group 17. _____

5. This Group 2 element has fewer protons than bromine, but more protons than sulfur. _____

6. This element has the lowest atomic number of any Group 16 element. _____

7. This element has an atomic number that is double the atomic number of silicon. _____

8. This element has more valence electrons than oxygen, fewer valence electrons than neon, more protons than sodium, but fewer protons than argon. _____

9. This element has an atomic number lower than that of aluminum and one less valence electron than the Group 16 elements. _____

10. This element is in Group 1 and has a higher atomic number than chlorine, but a lower atomic number than bromine. _____

Interpreting the Periodic Table

Examine the hypothetical periodic table shown below, and use it to answer the questions that follow.

K												
										J		
F												
			D			**R**						

1. Which pair of elements has the same number of valence electrons?_____

2. Which pair of elements is in the same period? _____

3. Which pair of elements is in the same family?
_____

4. Which element has the smallest atomic number? _____

5. If the atomic number of Element D is 20, then what is the atomic number of element R? _____